



Direct decomposition of methane over Ni catalyst supported in magnesium aluminate

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ARTICLE INFO

Article history:

Received 13 December 2011

Received in revised form 11 February 2012

Accepted 13 February 2012

Available online 21 February 2012

Keywords:

Methane decomposition

Magnesium aluminate

MgAl₂O₄

Synthesis

ABSTRACT

In this paper, the results obtained in the methane decomposition reaction promoted by a Ni catalyst supported on MgAl₂O₄ spinel are presented. The textural properties of the catalyst were investigated by X-ray diffraction (XRD), N₂ adsorption/desorption isotherms (BET and BJH methods), thermogravimetric (TGA) and temperature-programmed reduction (TPR) analysis. The influence of the operating conditions employed on the methane decomposition was studied. According to the results, it was found that a N₂:CH₄ molar ratio of 7:1 and catalyst reduction temperature of 700 °C/1 h were the best conditions for methane decomposition. In addition, in the methane decomposition over Ni/MgAl₂O₄, multiwall carbon nanotubes were formed.

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1. Introduction

Methane decomposition is a moderately endothermic reaction and, in general, only hydrogen is detected as the gaseous product. This process has been receiving increasing attention as an alternative route for the production of CO/CO₂-free hydrogen [1,2]. Unlike the conventional methods of hydrogen production from natural gas, *i.e.*, methane steam reforming and methane partial oxidation, which produce a mixture of hydrogen and carbon oxides, the catalytic cracking produces hydrogen and solid carbon, thereby eliminating the necessity for the separation of hydrogen from the other gaseous products. This significantly simplifies the process, reduces overall CO₂ emissions and makes the method particularly attractive for fuel cell applications, where hydrogen is the preferred fuel [3,4]. In addition, the decomposition of methane also results in the generation of a very important by product, nanocarbon materials (carbon nanotubes and nanofibers), that have been gaining considerable attention due to their excellent properties and potential applications [3–5].

Concerning the catalysts for the decomposition of methane, it is well known that supported Ni is one of the most effective

catalysts [1,6]. This metal has been supported over different materials (MgO, Al₂O₃, SiO₂, TiO₂, ZrO₂, MgO-SiO₂, H⁺-ZSM-5, USY, MCM-22, MCM-41, NiCuAl, LaNiO₃) and the effect these supports in the decomposition methane reaction has been examined [7–13].

Magnesium aluminate spinel, MgAl₂O₄, is an important ceramic and refractory material with several applications in many industries due to its properties, such as high melting point, low thermal expansion, high resistance to chemical attack, good mechanical strength, low dielectric constant and excellent optical properties [14–17]. However, the use of this material as a support for metallic catalysts is a relatively new application that has provided good results due to some of its attractive properties, such as low acidity, hydrophobic character, high thermal resistance and good interaction with the metallic phase, which are of interest for catalytic purposes [18–20]. We have recently shown that the magnesium aluminate spinel (MgAl₂O₄) could be used as support of metals for the hydrogen production by direct decomposition of methane [21]. However, the use of this material as metal support for this specific reaction is scarce in the literature.

In the present work, the performance of Ni catalyst supported on MgAl₂O₄ was investigated in direct decomposition of methane. The influence of the operating conditions N₂:CH₄ molar ratio and catalyst reduction temperature employed on the methane decomposition was investigated. In addition, structure of carbons formed by the reaction was analyzed.

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2. Experimental

2.1. Catalyst preparation

The MgAl_2O_4 support was prepared as described by Foletto et al. [19]. In the preparation of the precursor, the following reactants were utilized: aluminum isopropoxide $[(\text{CH}_3)_2\text{CHO}]_3\text{Al}$ and magnesium ethoxide $(\text{C}_2\text{H}_5\text{O})_2\text{Mg}$, both from Aldrich (purity >98%). This material was utilized as support in the preparation of nickel catalyst. For the precursor synthesis, two solutions 0.2 M of metallic alkoxides were prepared using alcohols as solvent. The first solution was prepared dissolving 62.8 g of aluminum isopropoxide (0.308 mol) in 1540 mL of isopropanol and the second one was prepared dissolving 12.56 g of magnesium ethoxide (0.110 mol) in 550 mL of methanol. Both solutions were mixed in a 5 L beaker then heated to methanol boiling point and under vigorous stirring it was added 330 mL of water. After adding water, the system was kept under heating and stirring for 3 h to complete the hydrolysis reaction. The formed material was separated from the alcohols by filtration and further dried in an oven at 120 °C for 24 h, resulting in a fine powder. This fine powder (mixture of Mg and Al hydroxides) was put in a 100 mesh sieve and calcinated in oxidizing atmosphere (air) at the temperatures of 700 °C, for 4 h, to form the spinel phase.

To obtain the Ni catalyst supported on MgAl_2O_4 , nickel was impregnated through successive wet impregnations with aqueous solutions of nickel nitrate hexahydrate, where the nominal content of nickel in the catalyst was kept at 20% (in weight). The impregnation was carried out at 80 °C. Then, samples were dried in an oven for 4 h, followed by calcinations at 700 °C for 5 h in air.

2.2. Catalytic tests

The decomposition reaction of CH_4 was carried out in a quartz-tube fixed bed flow reactor (9 mm i.d.) heated by an electric furnace. Before reactions, the catalysts (100 mg) were reduced *in situ* in an H_2 stream at 550 °C/1 h and 3 h and 700 °C/1 h and 3 h, with a heating rate of 10 °C min^{-1} . The experiments were conducted under atmospheric pressure at 550 °C. The reaction gas was composed of N_2 and CH_4 at molar ratios of 1:1; 1:3; and 7:1 ($\text{N}_2:\text{CH}_4$). The total flow rate of the reaction gas was 80 mL min^{-1} . The reactant and the product gases were analyzed with a Shimadzu GC-8A gas chromatograph, equipped with a thermal conductivity detector, a Porapak-Q column and a 5A molecular sieve column with Ar as the carrier gas. The N_2 in the reaction gas was used as a diluent and as an internal analysis standard. In this study, the catalytic activity was evaluated in terms of methane conversion.

2.3. Catalysts characterization

The X-ray diffraction (XRD) measurements were taken with a PanAnalytical X'pert PRO Multi-Purpose Diffractometer using $\text{Cu K}\alpha$ radiation ($\lambda = 1.5418 \text{ \AA}$) operating at 45 kV and 40 mA. The structural characterization obtained from the XRD data, including the average crystallite size and microstrain determination, were carried out using the GSAS+EXPEGUI program package [22,23], crystallographic information contained in the ICSD database [24] and the Rietveld method. Both the size and microstrain (lattice distortions) of a crystal cause XRD line broadening, and thus the profile function used in this study (number 4 in GSAS+EXPEGUI package) was a convolution of a pseudo-Voigt and asymmetry function, which also uses the microstrain broadening description of Stephens [25]. The XRD pattern line-widths were used to obtain the average size of the crystallites (L) and the microstrain (σ) using the formalism shown in Stephens [25], taking into account the instrumental broadening measured with a Y_2O_3 standard sample. In each case, background parameters, phase fractions, peak shape

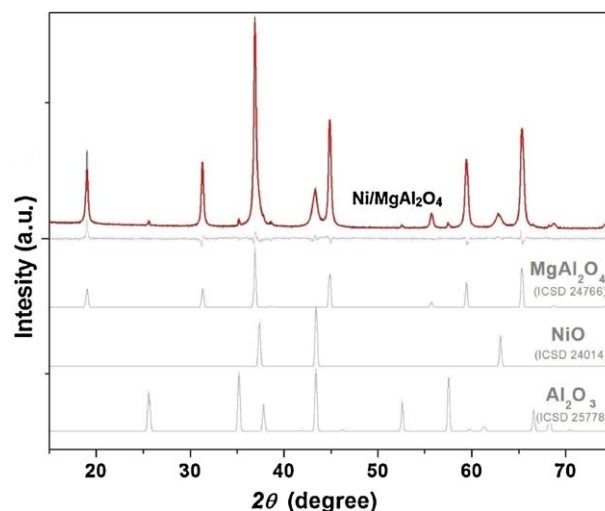


Fig. 1. X-ray diffraction patterns for the Ni/MgAl₂O₄ catalyst. Experimental and calculated (red line) X-ray diffraction patterns for the catalyst. The light gray lines represent the difference between experimental and calculated patterns and dummy patterns of the cubic MgAl₂O₄, cubic NiO and rhombohedral Al₂O₃ phases. (For interpretation of the references to color in this figure legend, the reader is referred to the web version of this article.)

parameters, cell parameters and sample position shift were refined before variation of further structural parameters. Isotropic dislocation (thermal) parameters were refined for all the atoms.

Samples were characterized by N_2 adsorption/desorption isotherms obtained at the temperature of liquid nitrogen in an automated physisorption instrument (Autosorb-1C, Quantachrome Instruments). Prior to the measurement, the samples were out-gassed under vacuum at 200 °C for 2 h. Specific surface areas were calculated according to the Brunauer–Emmett–Teller (BET) method and the pore size distributions were obtained according to the Barret–Joyner–Halenda (BJH) method.

Temperature-programmed reduction (TPR) analysis was performed to determine the reducible species present at the surface of the catalyst and the temperature at which these species are reduced by H_2 consumption. TPR was carried out with Micromeritics Chemisorb 2705 equipment, using 50 mg of catalyst and a temperature ramp from 25 to 1000 °C at 10 °C min^{-1} . A 30 mL min^{-1} flow rate of 5% H_2/N_2 was used.

The thermogravimetric analysis (TGA) was carried out in SDT 2960 thermobalance TA Instruments analyzer in Pt crucibles. The analysis was carried out under airflow of 50 mL min^{-1} on heating from room temperature to approximately 900 °C with a heating rate of 10 °C min^{-1} and at the latter it was carried out under airflow of 1.67 mL min^{-1} on heating from room temperature to approximately 900 °C with a heating rate of 5 °C min^{-1} .

The morphology of the carbon-containing materials was examined by scanning electron microscopy (SEM) using a Philips XL30 scanning microscope operating at an accelerating voltage of 20 kV. Transmission electron microscopy (TEM), with a JEM-1011 microscope operating at an accelerating voltage of 120 kV and Raman spectroscopy obtained on a Raman spectrometer (Renishaw – RG22) with an Ar laser at a wavelength of 514.5 nm.

3. Results and discussion

3.1. Characteristics of Ni/MgAl₂O₄

Fig. 1 shows the X-ray diffraction (XRD) pattern of the Ni/MgAl₂O₄ sample, the pattern calculated (using the Rietveld method) and three dummy patterns of the cubic MgAl₂O₄, cubic

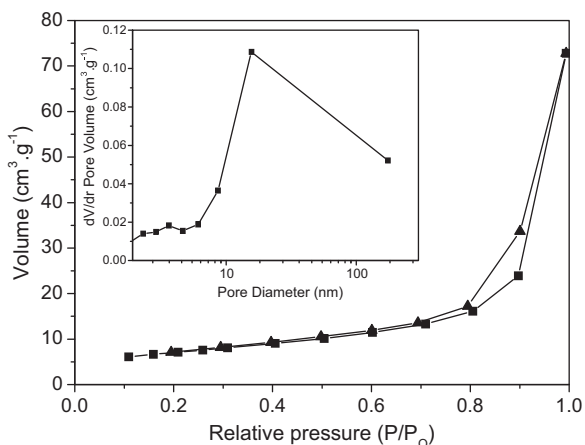


Fig. 2. N_2 adsorption/desorption isotherms and pore diameter distribution for the Ni/MgAl₂O₄ sample.

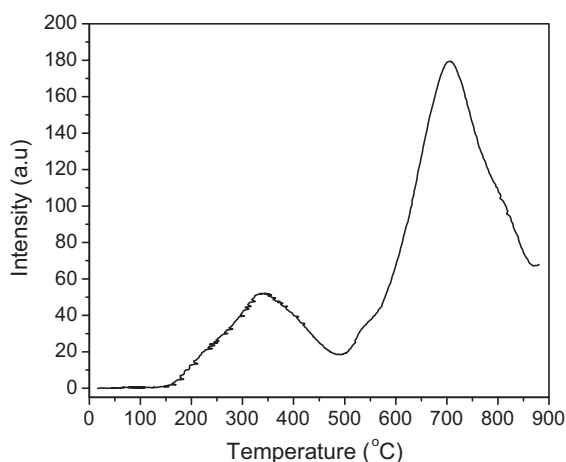


Fig. 3. TPR profile of Ni/MgAl₂O₄ catalyst.

NiO and rhombohedral Al₂O₃ phases. It is important to note that NiAl₂O₄ and MgAl₂O₄ spinel-like phases are indistinguishable in XRD analysis. The general AB₂O₄ spinel structure consists of A²⁺ and B³⁺ ions where the cations can easily displace each other. Therefore, Mg²⁺ in MgAl₂O₄ substitutes Ni²⁺ upon heat treatment at 700 °C in oxidizing atmosphere (air). NiAl₂O₄ spinel is created during these processes [26,27]. The Rietveld analysis was carried out based on this information and indicated 85 wt.% of MgAl₂O₄ with a lattice parameter $a = 8.08035 \text{ \AA}$, 12 wt.% of NiO with $a = 4.18177 \text{ \AA}$ and 3 wt.% of Al₂O₃ with $a = 4.76217 \text{ \AA}$ and $c = 13.0016 \text{ \AA}$. The mean crystallite sizes (microstrains) derived from the Rietveld results (peak line widths) were 390 Å (0.4%) for MgAl₂O₄, 150 Å (1.0%) for NiO and 1400 Å (0.3%) for Al₂O₃. The value obtained for the spinel MgAl₂O₄ phase agrees with the value of 342 Å for the sample produced from the mixture of aluminum and magnesium hydroxides and calcined at 1100 °C [19].

Fig. 2 shows the N_2 adsorption/desorption isotherms and pore size distribution of the Ni/MgAl₂O₄ sample. The curves presented in Fig. 2 are type-III (IUPAC). This type of isotherm has the characteristics of a system where the adsorbate molecules have a stronger interaction with each other than with the solid [28]. The profile of the pore size distribution (see insert in Fig. 2) indicates the predominance of mesopores. The surface area was 26.2 m² g⁻¹ and pore volume of 0.126 cm³ g⁻¹.

Fig. 3 shows the temperature-programmed reduction (TPR) curve for the Ni/MgAl₂O₄ sample, where two peaks can be observed

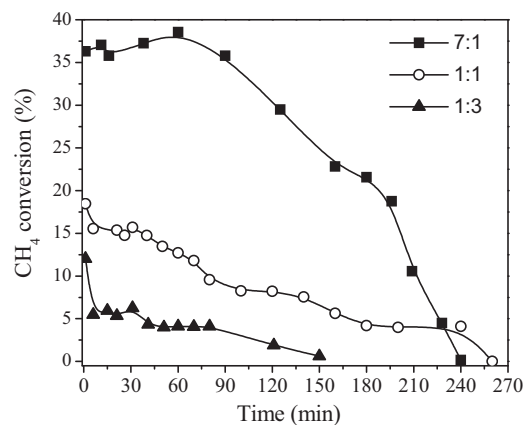


Fig. 4. Methane decomposition over Ni/MgAl₂O₄ catalyst at 550 °C with different $N_2:CH_4$ molar ratios. Reduction step: 700 °C/1 h.

at around 350 °C and 700 °C. The first peak can be attributed to NiO reduction weakly interacted with MgAl₂O₄ and the second peak to the reduction of Ni and Mg strongly interacted with MgAl₂O₄ spinel [26,29], corroborating the XRD results.

3.2. Methane decomposition over Ni/MgAl₂O₄ catalyst

In order to investigate the catalytic activity of the Ni/MgAl₂O₄ catalyst, the methane decomposition reaction was carried out. Fig. 4 shows the catalyst behavior with different $N_2:CH_4$ molar ratios as a function of the reaction time.

The results showed an increase in the average methane conversion values with an increase in the $N_2:CH_4$ molar ratio. Under a feed flow condition with the methane most diluted the catalyst gave the highest average conversion value (31%) and the highest initial catalytic activity (37%). On the other hand, when the methane was very concentrated in the reaction feed flow, that is, using $N_2:CH_4$ in molar ratios of 1:1 and 1:3, the catalyst gave lower conversion values, not exceeding 18%. This behavior may be due to the amount of active sites present in this catalyst not being sufficient to convert all of the reactive molecules in contact with the catalytic surface because of the increase in the reaction feed flow mixture concentration. Thus, a given amount of reagent would pass through the catalytic bed without undergoing any kind of effective interaction with the catalytic surface, remaining at higher concentrations in the reaction effluent, causing lower methane conversion values. Moreover, in all tests, an ongoing catalytic deactivation until a total loss

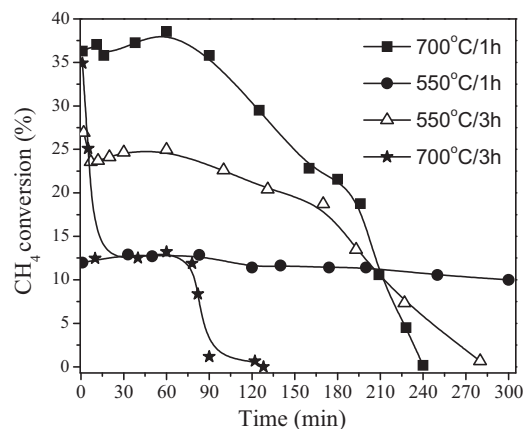


Fig. 5. Methane decomposition over Ni/MgAl₂O₄ catalyst at 550 °C and $N_2:CH_4$ molar ratio of 7:1, with different conditions of reduction step.

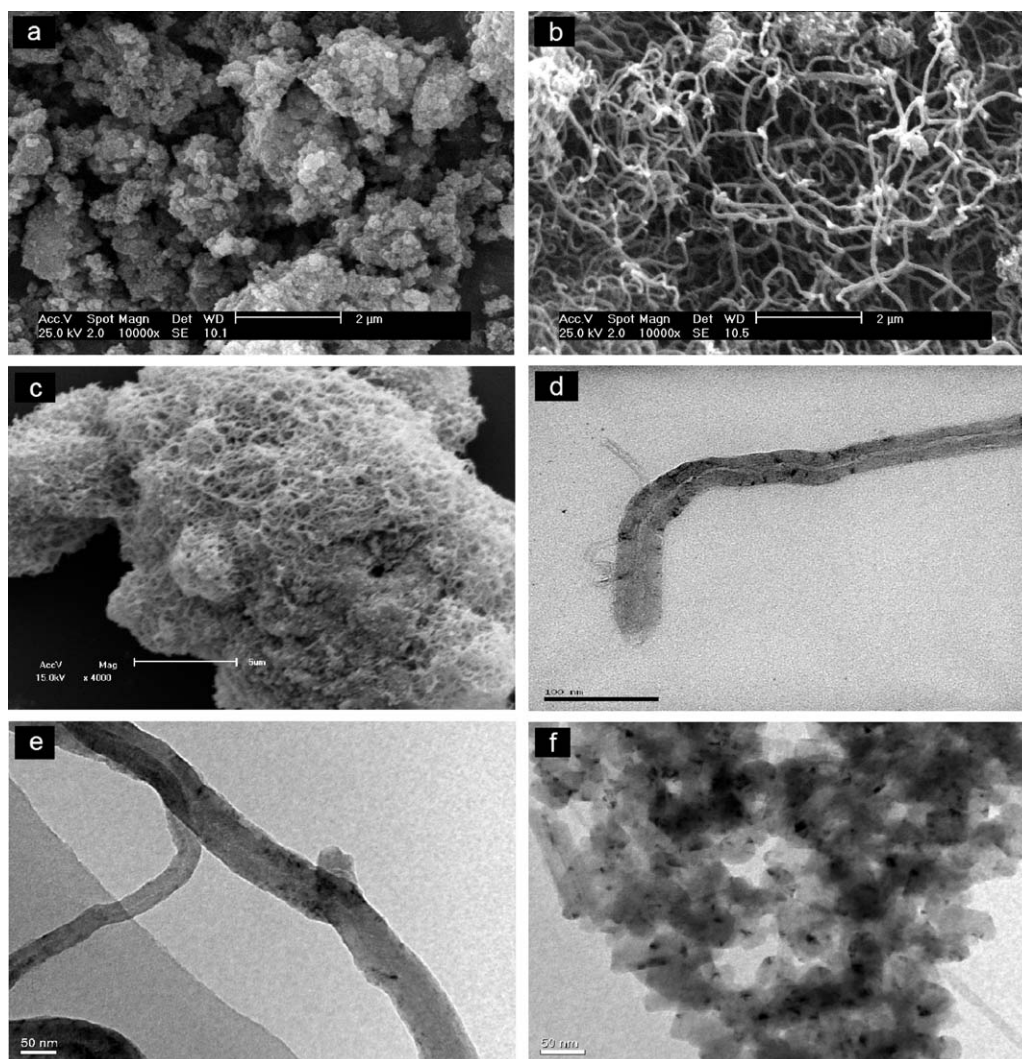


Fig. 6. SEM images of the NiMgAl₂O₄ catalyst surfaces (a) before, and (b) N₂:CH₄ = 7:1, reduced at 700 °C/1 h, (c) N₂:CH₄ = 7:1, reduced at 550/1 h and TEM images of the same catalyst (d) N₂:CH₄ = 7:1, reduced at 700 °C/1 h, (e) and (f) N₂:CH₄ = 1:3, reduced at 700 °C/1 h. Reaction temperature: 550 °C.

of activity, at 240 min of reaction for a N₂:CH₄ molar ratio of 7:1, at 260 min of reaction for a N₂:CH₄ molar ratio of 1:1 and at 150 min for reaction of a N₂:CH₄ molar ratio of 1:3, was observed. However, the catalyst with a N₂:CH₄ molar ratio of 1:3, presented the lowest catalytic activity, not exceeding 12%. This is mainly due to the more extreme working conditions where the formation of carbonaceous materials is favored, giving rise to the surface blocking of the active site [30]. During the catalytic decomposition reaction, methane molecules are initially adsorbed (with dissociative adsorption) and decomposed on the metal surface of the catalyst particle, resulting in the formation of chemisorbed carbon species and the release of gaseous hydrogen; the carbon species then dissolve and diffuse through the bulk of the metal particle. Deactivation occurs when the rate of carbon diffusion through the metal catalyst particle is slower than the rate of carbon formation at the surface of the active metallic sites. Under these circumstances, carbon builds up at the catalyst surface and eventually encapsulates the metal particle causing activity loss [30,31].

Fig. 5 shows the catalytic behavior of the Ni/MgAl₂O₄ catalyst with different reduction temperatures and times as a function of the reaction time.

In all of the tests, catalytic deactivation was observed, except when the catalyst was submitted to the reduction process at

550 °C/1 h. According to the results, it can be observed that the tests carried out with the activated catalyst at 700 °C/1 h gave higher average values for the CH₄ conversion (31%) and a greater initial catalytic activity (37%) compared with the catalyst which underwent reduction at 550 °C/1 h and 550 °C/3 h. The higher reduction temperature may have favored the reduction in reducible species leaving the active phase more available for catalytic activity. The catalyst reduced at 700 °C/3 h showed a rapid and gradual decrease in catalytic activity, with lower average methane conversion values (11%) and shorter reaction time, not exceeding 130 min. This finding may be related to the possibility that this catalyst underwent sintering due to the activation time and temperature, which leads to the active phase being less accessible to the flow of reagent which passes through the reactor bed and thus not converting these reagent molecules into carbon and hydrogen products. The catalyst reduced at 550 °C/1 h presented an average methane conversion of 12%, which remained almost constant until the end of the reaction (300 min). It is probable that the reduction condition employed was not sufficient for the effective activation of the catalytic sites. The catalyst reduced at 550 °C/3 h presented initial catalytic activity of around 24% of the CH₄ conversion and total loss of activity at 280 min of reaction.

Table 1Data of mass loss of the Ni/MgAl₂O₄ sample after catalytic decomposition of the methane as a function of the operational conditions.

Entry	Operational conditions			TGA	
	N ₂ :CH ₄	T _R ^a (°C)	T _{Red} ^a (°C)/time (h)	Mass loss (%)	Temperature (°C)
1	1:3	550	700/1	66	625
2	1:1	550	700/1	86	586
3	7:1	550	700/1	3–66	250–600
4	7:1	550	700/3	4	546
5	7:1	550	550/1	55	625
6	7:1	550	550/3	79	607

^a T_R, reaction temperature; T_{Red}, reduction temperature.

3.3. Characterization of deposited carbons

TGA is a very powerful technique for determining the quality of the synthesized CNTs due to differences in oxidation stability and degree of graphitization. It appears that various types of carbon such as amorphous carbon and CNT possess different oxidation temperatures and can be distinguished using the TGA technique [32]. Results of mass loss determined from the thermogravimetric curves (not shown here) of the Ni/MgAl₂O₄ catalyst after the catalytic tests under all operational conditions studied in this work are shown in Table 1.

The thermogravimetric analysis of the Ni/MgAl₂O₄ catalyst after the catalytic tests at 550 °C showed a mass loss in the region of 546–625 °C. It can be observed for entries 1, 3, 5 and 6 that there was a mass loss of 66% at 625 °C and 600 °C, 55% at 625 °C and 79% at 607 °C, respectively. The mass loss of this material at above 600 °C can be attributed to the MWNTs [31,33]. It can also be noted from entry 3 that there was a mass loss of 3% at 250 °C, which can be attributed to amorphous carbon which decomposes at temperatures below this value. In the case of entries 2 and 4, there was a mass loss of 86% at 586 °C and 4% at 546 °C. This may be associated with small amounts of amorphous particles in the nanotubes, the influence of metal particles and/or defects on the surface of the nanotubes. These factors can affect the temperature at which the maximum decomposition rate occurs. It can be concluded that in any of the operational conditions employed, the formation of material more steady as the CNTs occurs [34–37].

The SEM and TEM characterizations of the carbon-containing materials after the CH₄ reaction are shown in Fig. 6. The SEM image presented in Fig. 6a shows the Ni/MgAl₂O₄ sample before of the catalyst tests. The images obtained in the SEM (Fig. 6b and c) and TEM (Fig. 6d–f) analyses show that after the catalytic tests there were modifications on the catalyst surface due to the deposition of carbonaceous material. These images suggest that different carbon species were formed according to the feed flow of the reaction and variations in the reduction temperature and time, with a significant formation of CNTs. The formation of carbon nanotubes with a diameter of approximately 40 nm can be clearly observed in Fig. 6d. It was identified that besides the deposition of carbon in the form of CNTs there was the presence of fibers and irregular agglomerations, as observed in Fig. 6e and f, respectively. The results show that there was the preferential formation of carbon nanotubes in the case of the Ni/MgAl₂O₄ catalyst in the decomposition reaction with higher molar ratios (N₂:CH₄ = 7:1). This result may explain, or at least be related to, the greater stability of the catalyst under this experimental condition. For the catalyst in the decomposition reaction with lower molar ratios (N₂:CH₄ = 1:3), besides the nanotubes, the formation of carbon agglomerations and nanofibers was identified, which may favor catalytic deactivation. The images in Fig. 6d–f show some black points which may be attributed to the nanoparticles of Ni and/or amorphous carbon, which were involved and deposited on the walls of carbon nanotubes. Similar behavior has been described by Hsieh et al. [38], Zhou et al. [39] and Guevara

et al. [40], employing metal and bimetal catalysts of Pt, Pt–Ni(Fe, Co) and Ni–Ce, respectively.

4. Conclusions

The results indicated that the Ni/MgAl₂O₄ sample is composed by nanometric crystallites of MgAl₂O₄ spinel phase (85 wt.%), NiO phase (12 wt.%) and Al₂O₃ (3 wt.%). It was observed that the catalytic performance, activity and stability, are dependent on the operational conditions employed. According to the results, the best conditions for methane decomposition were N₂:CH₄ molar ratio of 7:1 and reduced temperature of 700 °C/1 h, at reaction temperature of 550 °C. The results here obtained also show that the Ni/MgAl₂O₄ catalyst is effective for the production of carbon nanotubes.

Acknowledgments

The authors are grateful to Federal University of Santa Catarina (UFSC) for infrastructure (including LCME, LDRX and LabMat facilities), and CNPq (Brazil) for financial support.

References

- [1] T.V. Choudhary, C. Sivadinarayana, A. Chusuei, A. Klinghoffer, D.W. Goodman, *J. Catal.* 199 (2001) 9–18.
- [2] N. Muradov, Z. Chen, F. Smith, *Int. J. Hydrogen Energy* 30 (2005) 1149–1158.
- [3] N.Z. Muradov, *Energy Fuels* 12 (1998) 41–48.
- [4] J.L. Pinilla, I. Suelves, R. Utrilla, M.E. Gálvez, M.J. Lázaro, R. Moliner, *J. Power Sources* 169 (2007) 103–109.
- [5] J.L. Pinilla, I. Suelves, M.J. Lázaro, R. Moliner, J.M. Palacios, *Int. J. Hydrogen Energy* 35 (2010) 9801–9809.
- [6] K. Otsuka, S. Kobayashi, S. Takenaka, *Appl. Catal. A: Gen.* 190 (2000) 261–268.
- [7] J. Chen, Y. Qiao, Y. Li, *Appl. Catal. A: Gen.* 337 (2008) 148–154.
- [8] J. Ashok, S.N. Kumar, A. Venugopal, V.D. Kumari, M. Subrahmanyam, *J. Power Sources* 164 (2007) 809–814.
- [9] S.K. Saraswat, S.S. Pant, *Int. J. Hydrogen Energy* 36 (2011) 13352–13360.
- [10] I. Suelves, J.L. Pinilla, M.J. Lázaro, J.M. Palacios, *J. Power Sources* 192 (2009) 35–42.
- [11] J. Ashok, G. Raju, P.S. Reddy, M. Subrahmanyam, A. Venugopal, *Int. J. Hydrogen Energy* 33 (2008) 4809–4818.
- [12] M.E. Rivas, C.E. Hori, J.L.G. Fierro, M.R. Goldwasser, A.G. Constant, *J. Power Sources* 184 (2008) 265–275.
- [13] S. Takenaka, H. Ogihara, I. Yamanaka, K. Otsuka, *Appl. Catal. A* 217 (2001) 101–110.
- [14] W. Kim, F. Saito, *Powder Technol.* 113 (2000) 109–113.
- [15] J.G. Li, T. Ikegami, J.H. Lee, T. Mori, Y. Yajima, *Ceram. Int.* 27 (2001) 481–489.
- [16] L.R. Ping, A.M. Azad, T.W. Dung, *Mater. Res. Bull.* 36 (2001) 1417–1430.
- [17] A.D. Mazzoni, M.A. Sainz, A. Caballero, E.F. Aglietti, *Mater. Chem. Phys.* 78 (2003) 30–37.
- [18] A.S. Bocanegra, A.D. Ballarini, O.A. Scelza, S.R. Miguel, *Mater. Chem. Phys.* 111 (2008) 534–541.
- [19] E. Foletto, R.W. Alves, S.L. Jahn, *J. Power Sources* 161 (2006) 531–534.
- [20] J. Guo, H. Lou, H. Zhao, D. Chai, X. Zheng, *Appl. Catal. A: Gen.* 273 (2004) 75–82.
- [21] G.D.B. Nuernberg, H.V. Fajardo, E.L. Foletto, S.M.H. Probst, N.L.V. Carreño, L.F.D. Probst, J. Barraut, *J. Catal. Today* 176 (2011) 465–469.
- [22] A.C. Larson, R.B.V. Dreele, Los Alamos National Laboratory Report LAUR 86-748, 2004.
- [23] B.H. Toby, *J. Appl. Cryst.* 34 (2001) 210–213.
- [24] Inorganic Crystal Structure Database (ICSD), Gmelin-Institut für Anorganische Chemie und Fachinformationszentrum, FlZ, Karlsruhe, 2007.
- [25] P.W. Stephens, *Appl. Cryst.* 21 (1999) 281–289.
- [26] D.S. Park, Z. Li, H. Devianto, H. Lee, *Int. J. Hydrogen Energy* 35 (2010) 5673–5680.

- [27] A. Monzón, N. Latorre, T. Ubieta, C. Royo, E. Romeo, J.I. Villacampa, L. Dussault, J.C. Dupin, C. Guimon, M. Montieux, *Catal. Today* 116 (2006) 264–270.
- [28] V.G. Teixeira, F.M.B. Coutinho, A.S. Gomes, *Quím. Nova* 24 (2001) 808–818.
- [29] H. Ozdemir, M.A.F. Oksuzomer, M.A. Gurkaynak, *Int. J. Hydrogen Energy* 35 (2010) 12147–12160.
- [30] G.B. Nuernberg, H.V. Fajardo, D.Z. Mezalira, T.J. Casarin, L.F.D. Probst, N.L.V. Carreño, *Fuel* 87 (2008) 1698–1704.
- [31] R.M. Almeida, H.V. Fajardo, D.Z. Mezalira, G.B. Nuernberg, L.K. Noda, L.F.D. Probst, N.L.V. Carreño, *J. Mol. Catal. A: Chem.* 259 (2006) 328–335.
- [32] C.M. Chen, Y.M. Dai, J.G. Huang, J.M. Jehng, *Carbon* 44 (2006) 1808–1820.
- [33] G.S. Gallego, J. Barrault, C.B. Dupeyrat, F. Mondragón, *Catal. Today* 149 (2010) 365–371.
- [34] A.B. Suriani, A.A. Azira, S.F. Nik, R.M. Nor, M. Rusop, *Mater. Lett.* 63 (2009) 2704–2706.
- [35] N. Das, A. Dalai, J.S.S. Mohammadzadeh, J. Adjaye, *Carbon* 44 (2006) 2236–2245.
- [36] E. Maccallini, T. Tsoufis, A. Policicchio, S.L. Rosa, T. Caruso, G. Chiarello, E. Colavita, V. Formoso, D. Gournis, R.G. Agostino, *Carbon* 48 (2010) 3434–3445.
- [37] P.Z. Poor, A. Badieli, A.A. Yousefi, B.D. Fahlman, A. Abbasi, *Catal. Today* 150 (2010) 100–106.
- [38] C. Hsieh, J. Lin, J. Wei, *Int. J. Hydrogen Energy* 34 (2009) 685–693.
- [39] M. Zhou, G. Lin, H. Zhang, *Chin. J. Catal.* 28 (2007) 210–216.
- [40] J.C. Guevara, J.A. Wang, L.F. Chen, M.A. Valenzuela, P. Salas, A. García-Ruiz, J.A. Toledo, M.A. Cortes-Jácome, C. Angeles-Chavez, O. Novaro, *Int. J. Hydrogen Energy* 35 (2010) 3509–3521.